**Name:**

**

*Rickard & Geelan 2008: Science Ways 3*

**Atoms and Compounds**

Complete the following tasks while viewing the PowerPoint presentation on Atoms and Compounds.

1. What is Warfarin used for? (2 uses)

1. Complete: Atoms are tiny that make up all .
2. Name the three main parts of an atom: 1. 2. 3.
3. Complete: The of the atom is in atom. It

is made up of the atom’s and . The move

around the at .

1. Complete the following table:

|  |  |
| --- | --- |
| **Atomic Particle** | **Electric Charge** |
| P\_\_\_\_\_\_\_\_\_\_\_\_\_ |  |
|  | No charge (neutral) |
| E\_\_\_\_\_\_\_\_\_\_\_\_ |  |

1. Why do atoms usually have a zero electric charge overall?

1. Complete: Ions are
2. If an atom gains an electron, what overall charge will it have?
3. If an atom loses an electron, what overall charge will it have?
4. How is the atomic number determined?
5. What is the atomic number of the carbon atom on the PowerPoint?
6. How is the mass number calculated?
7. What is the mass number of the carbon atom on the PowerPoint?
8. An element is a substance that is made up of
9. List 4 examples of an element: 1. 2. 3. 4.
10. How is a compound different to an element?

What is the difference between water and hydrogen peroxide?

**Review Notes: Make sentences form each of the following sentence starters**

An Atom is;

An atom is made up of;

The number of protons in an atom is its;

The number of protons plus neutrons in an atom is its;

An ion is;

The two types of ions are;

Positive ions are formed when;

Negative ions are formed when;

Metal atoms form;

Non-metal atoms form;

An element is;

A compound is;